

Warmup - pH + pOH

Name _____

Determine pH

1. $[H^+] = 1E-3$

2. $[H^+] = 1E-14$

3. $[H^+] = 1E-7$

4. $[H^+] = 1E-10$

Determine pOH

5. $[OH^-] = 1E-8$

6. $[OH^-] = 1E-2$

7. $[OH^-] = 1E-12$

Determine pH

8. $[OH^-] = 1E-13$

9. $[OH^-] = 1E-2$

10. $[OH^-] = 1 \times 10^{-6}$

11. $[H^+] = 10^{-11}$

Determine pOH

12. $[H^+] = 1E-5$

13. $[OH^-] = 1E-12$

14. $[H^+] = 1E-14$

15. $[H^+] = 1E-7$

Determine $[H^+]$

16. $pH = 3$

17. $pH = 11$

18. $pH = 7$

Determine $[OH^-]$

19. $pOH = 2$

20. $pOH = 14$

21. $pOH = 8$

Determine $[H^+]$

22. $pOH = 13$

23. $pOH = 2$

24. $pH = 6$

25. $pOH = 9$

Determine $[OH^-]$

26. $pH = 4$

27. $pOH = 6$

28. $pH = 12$

29. In a solution, $[H^+] = 1 \times 10^{-4}$. What is $[OH^-]$?

30. In a solution, $[OH^-] = 1 \times 10^{-2}$. What is $[H^+]$?

31. The pH of a solution is 6. What is its pOH ?

32. The pOH of a solution is 12. What is its pH ?

33. A solution has a $pH = 4$. Is it acidic or basic?

34. A solution has a $pOH = 2$. Is it acidic or basic?

35. What is an indicator?

36. Is an acid a proton (H^+) Donor or a proton (H^+) Acceptor?