**More Thermochemistry Problems (L1)**

1. How many joules are needed to warm 675 grams of water from 12.00C to 85.00C?

2. How much energy is needed to melt 35g of ice at 00C?

3. How much energy is released when 5.0g of water cool from 750C to 250C?

1. What was the mass of a sample of water if the addition of 4.28 x 103 joules raised the temperature of the sample from 22.00C to 34.50C?

1. If 750J of heat energy were released from condensing steam, what mass of steam condensed?
2. How many joules are needed to raise the temperature of 500. g of chromium, specific heat 0.448 J/g0C, from 26.00C to 95.00C?

1. What is the specific heat of a sample of an alloy if 3.75 x 103 joules of heat are released when a 315 gram sample of alloy cools from 78.00C to 28.40C?

1. What will be the final temperature if 3.50 x 104 joules of heat are added to a 454 gram sample of iron at 24.00C?

1. Calculate the number of joules needed to warm 275 grams of copper from 25.00C to 9500C.

1. What will be the final temperature when 8.75 x 103 joules are added to 75.0 grams of water at 23.00C?

1. When 3.88 x 104 joules were added to a sample of iron, the temperature rose from 24.50C to 1780C. What was the mass of the iron sample?